If it took 652seconds for a sample of B_2H_6 to effuse, for which of the following would it require 823seconds for the same number of moles to effuse?

○ CO_2 ○ H_2 ○ HCN ○ Xe ○ Ne

In a given period of time, 0.778moles of B_2H_6 effuses. How many moles of NH_3 would effuse in to the same period of time?

If 0.0142 moles of C_2H_2 effuses in 775seconds, how many seconds would it take for the same number of moles of CO to effuse?

[ ]

Graham's Law of Effusion Practice

Ne has an effusion rate of 0.00624mol/sec. Which of the gases below would have an effusion rate of 0.00701mol/sec?

○   Ar ○   H₂ ○   COCl₂ ○   CH₄ ○   He

If Xe has an effusion rate of 0.00669moles/sec. What would the effusion rate for CO be in moles/sec?

Which of the following would effuse at a rate equal to 31.5% of the rate of H₂?

○   CH₃Cl ○   Ne ○   HCN ○   B₂H₆ ○   CO₂